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1. Name- Cadet Ankan Chatterjee Roll Number- 4854 Standard- XII-B Project Name: Preparation ³ jab³ ³n and determination of foam capacity of different soaps Session: 2021-22 2. Sainik School Purulia purulia, Certificate of West Bengal Department of Qu'Amica is to certify that Ankan Chatterjee, a Class XII student, has successfully completed the research on the project mentioned below under the ³ of Mr. S.N Mishra (master of subjects) during the 2021-2022 period in the partial fulfillment of the Practical Examination of Qu'mica conducted by AISSCE , New Delhi. Signature of the signature (external examiner) (chemistry teacher) 3. Recognition in the achievement of this project with success, many people have bestowed their blessings and the heart ³ promised support on mA. A big thank you to everyone who has been concerned about this project. I would like to thank our master of chemistry, Mr. S.N Mishra, whose valuable guide has been those who helped me repair this project and have a complete success. His suggestion and instructions have served as the main contribution ³ completing the project. Then I would like to thank my parents and friends who have helped me with their valuable suggestions and guidance ³ various phases of the completion ³ the project. Last but not least, I would like to thank our lab assistant, se&±or Sudip Mahata. 4. Declaration: I hereby declare that the work of the project entitled "Preparation ³ jab³ ³n and determination of the foam capacity of different soaps", submitted to the Department of Qu'Amica, Sainik School Purulia for the chemistry of the subject under the guise of Mr. S.N Mishra is a record of the original work done by mA. Furthermore, I declare that this project or any part of it has not been presented elsewhere for any other class. 5. Content 1. Introduction³ 2. Experiments 3. Theory 4. Conclusion³n 5. Bibliography 6. ed ed olixordih ed sodavired omoc esraredisnoc nedeup selonacla sol o selohocla sol. ?sasarg sal y setieca sol nos ©AuQcÁ lareneg n'Aicamrofni PAOS ed ehT. sedirecylg dexim dellac era esehT .edirecylg emas eht ni tneserp si lacidar dica eno naht erom yllausU .9 .nitimlapirt ro nitimlap sa deman si lacidar dica citimlap eht gniniatnoc eno eht dna niraetsirt ro niraets dellac si lacidar dica cidar ciraeets gniniatnoc edirecylg eht .ecnatsni roF .meht ni tneserp lacidar dica eht ot seht segnidrocca deman era direcylg ehT .)HOOC53H71C (hint ciraeets dna)HOOC13H51C (hint citimlap .)HOOC72H31C (hint citsiryrm .)HOOC33H11HC(hcus sdica detarutas fo noitroporp regal a niatnoc staf eht gnimrof sedirecylg eht elihw .,cte .)HOOC13H71C (hint cieloniil . 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The capillary is saddled in a tube of oily hole in en in en. A bulb is blown and two marks are stamped on the stem of the tube, one above and one below the bulb. 12. (The foam capacity of that sample of jab³n be greater in which case it takes longer to disappear.) But all soaps are not equally effective in their cleaning ³ . ³ Soaps are the sodium or potassium salts of higher fatty acids such as palm acid (C15H31COOH), east acid (C17H35COOH), oleic acid (C17H33COOH), etc., its general formula is Rcoona or RCOOK where R is the longest chain group I.E.E.E.E.E.E.E.E.E. C15H3111 , /C17H33, etc., therefore, each mole of jab³n consists of two parts: a part ripofAlca (soluble in oil) and a part hydrA³fila (soluble in water) Coona or cook. The cleaning action ³ jab³n depends on the solubility of the long alkyl in fat or oil drops and -coona or -cook in the water. Dirt is kept on the fabric surface by grease or oil drops. Each time the jab³n is applied over a dirty paA±o hAmedo, the non-polar alkyl group dissolves in fat (non-polar), while the polar-crown group dissolves in water (polar). In this way, an emulsion is formed between fat ³ water that appears as foam. The dirt along with the emulsi³n is washed when the fabric is treated with excess water. Therefore, the washing capacity of the jab³n depends on its 13. Foaming capacity, i.e. the extent to which it foams with water. Foam capacity also depends on the quality of the water used. If soft water is used, soaps easily produce foam. On the other hand, if hard water is used, even good quality soaps will not produce foam. The root³n is that hard water contains magnesium and calcium ions that form insoluble magnesium and calcium salts of carboxylic acids that in turn precipitate in the form of slag. 2C17H35COONA + CA2 + Ae a a(C17H35COO) 2 CAe a a + 2NA + (WATER SOLUBLE) (PPT) 2C17H35COONA + MG2 + Ae a a MgAe a a + 2NA + (soluble water) (ppt.) However, If the hard water is first treated with NA2CO3, the magnesium and calcium ions are presented and precipitated as their insoluble carbonates. Filtering can be used for washing purposes as it is now made soft. In contrast, detergents can be used to wash purposes even in hard water. The reason ³ is that detergents are sodium or potassium salts of ³ or aromatic acids and even their calcium and magnesium salts are soluble in water and thus do not form slag, but rather form foam when treated with hard water. 14. Experiment- Objective to compare the foam capacity of different soaps. Apparatus five ³ flasks of 100 ml, five test tubes of 20 ml, Cylinder of measurement of 100 ml, support of test tube, weight box and stop clock. ³ . Chemicals are required: five different samples of ³ , distilled water. Theory: the foam capacity of the jab³n depends on the nature of the jab³n and its concentration³ This can be compared by shaking equal volumes of the solutions of different soap samples that have the same concentration ³ the same force and aa for the same period of time. Solutions are then allowed to stop when the foam during shaking gradually disappears. The time taken for foaming is determined on each sample. The longer it takes for the foam to disappear ³ the given sample of ³ , the greater its foam capacity or cleaning capacity. 15. Procedure: Ae& Take five 100 ml ³ flasks and give them as 1. 2. 3. 4 and 5. Place 50 mL of distilled water in each of the ³ flasks and add 5 g of jab³n chips or grains of different jab³n samples to each flask. Hot the contents of each flaskto get a clear solution³ Ae a Take five 20 ml test tubes and add 10 ml of distilled water to each of them and onu onu rop etnemasorogiv odinetnoc us etiga y raglup le noc 1 orem³An oyasne ed obut led acob al erreic eA ~a a .etneidnopserroc oyasne ed obut la ocin³Ac zartam adac ed n³Abaj ed n³Aiculus ed lm 1 eugerga aroha .5 y 4 .3 .2 .1 omoc Place test tube in the test tube stand and start the stop watch immediately. Note the time taken for disappearance of the foam produced. eAAeA Repeat the same procedure for test tubes 2, 3, 4 and 5 shaking each time with the same force and for the same time (one minute). Note the time taken for disappearance of foam and record the observations in a tabular form. OBSERVATION Amount of each soap sample = 5.0 gm Volume of distilled water added to each soap sample = 50.0ml 16. Test Tube Number Name of the Soap Sample Volume of Soap Solution Added Volume of Water Added Time of disappearance of Foam 1. Dove 1.0ml 10.0ml 2. Lifebuoy 1.0ml 10.0ml 3. Dettol 1.0ml 10.0ml 4. Lux 1.0ml 10.0ml 5. Liril 1.0ml 10.0ml 6. Medimix 1.0ml 10.0ml RESULT: The foaming capacity and hence the cleansing capacity of different samples of soaps is in the order: Lifebuoy > Dettol > Liril > Lux > Medimix > Dove PRECAUTION: Each test tube containing the soap solution must be shaken with the same force and for the same period of time. 17. BIBLIOGRAPHY eAAeA Textbook of Class 12 Chemistry eAAeA eAAeA eAAeA eAAeA Help from Subject Teacher and Lab Attender. Attender.

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